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88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g.

Calculate the percentage yield. 2. 6.0 mol of N₂

are mixed with 12.0 mol of H₂ according to the following equation: N

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Calculate the percentage

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yield. 2. 6.0 mol of N_2
are mixed with 12.0 mol
of H_2 according to the
following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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space provided. 1. Given the following equation:
$$\text{C}_3\text{H}_4(\text{g}) + x \cdot \text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$$

a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ?
2. a. What ...

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