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Section 6.2 Covalent Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. jada5466. Terms in this set (15) Covalent Bond is a. Is a chemical bond in which two atoms share valence electrons. Which ones show orbitals of atoms overlapping when a covalent bond forms? [electrons dot , structural formula , space ...

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Chapter 6 Chemical Bonds. Section 6.2 Covalent Bonding (pages 165 – 169) This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules. Reading Strategy (page 165) Relating Text and Visuals As you read the section, look closely at Figure 9.

~~Section 6.2 Covalent Bonding – Yumpu~~

Section 6.2 Covalent Bonding (pages 165 – 169) This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules. Reading Strategy (page 165) Relating Text and Visuals As you read the section, look closely at Figure 9.

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6.2 Covalent Bonding Covalent Bond - a chemical bond in which two atoms share a pair of valence electrons. o Nonmetals do not usually transfer electrons, so they generally form covalent bonds o Nonmetals share valence electrons – which is a covalent bond. o Four ways to represent a covalent bond. o Electron Dot diagram

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***Covalent bond in which electrons are not shared equally. This occurs when one atom has a higher electronegativity than the atom it is sharing with.-Whenever two different atoms form a covalent bond, a polar bond is formed-Atoms with a greater

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attraction for electrons has a partial negative charge, and the other atom has a partial positive charge

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Section 6.2 – Covalent Bonding A covalent bond is a chemical bond in which two atoms share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a single bond. Covalent vs Ionic Bond

~~Chapter 6 — Chemical Bonds~~

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~~Ch. 6 (Section 6.2 Workbook Questions), Chemical Bonds ...~~

Chemical Bonds 6.2 6.2 Covalent Bonding. Question Answer; In a _____ bond, the valence electrons are shared between two atoms. Covalent: What is a neutral group of atoms joined together by one or more covalent bonds? Molecule: True or false: Opposites charges repel.

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~~Free Physical Science Flashcards about Chemical Bonds 6.2~~

6.2.1 Describe how covalent bonds form and the attractions that keep atoms together in molecules. 6.2.2 Compare polar and nonpolar bonds, and demonstrate how polar bonds affect the polarity of a molecule. 6.2.3 Compare the attractions between polar and nonpolar molecules. Build Vocabulary Concept Map Have students construct a concept map using the terms atoms,

~~6.2 Covalent Bonding – Physical Science~~

covalent bonding, a bond forms from the sharing of electron pairs between two atoms. In a purely covalent bond, the shared electrons are “ owned ” equally by the bonded atoms. LOUISIANA STANDARDS LA.PS.24 Describe the influence of intermolecular forces on the physical and chemical properties of covalent compounds. (PS-H-C5) CHEMICAL Bonding 165

~~CHAPTER 6 Chemical Bonding~~

SECTION 2 Name Class Date Ionic and Covalent Bonding continued MULTIPLE BONDS Some atoms need to share more than one pair of electrons to fill their outermost energy levels. 2e – 4e – 4e – Oxygen 2e – 4e – 2e – 2e 6e – Nitrogen 2e 2e O O N N Double covalent bond Triple covalent bond Four electrons are in the shared electron cloud. Six ...

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~~CHAPTER SECTION 2 Ionic and Covalent Bonding~~

Section 2- Covalent Bonding and Ionic Compounds Objectives: define molecule; explain the difference between a chemical formula and a molecular formula; define bond energy; describe the octet rule; list the exceptions to the octet rule; write Lewis dot diagrams; write Lewis structures for molecules; list and describe multiple bonds; define resonance

~~Ch 6 – Honors Chem Wins~~

View Ch 6-Chemical Bonding section 1.pdf from CHEMISTRY 101 at Keiser University. a b a b c c a valence electrons ionic bond polar covalent ionic bond 1.7 H 2 HCl NaCl Shared pair of electrons

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